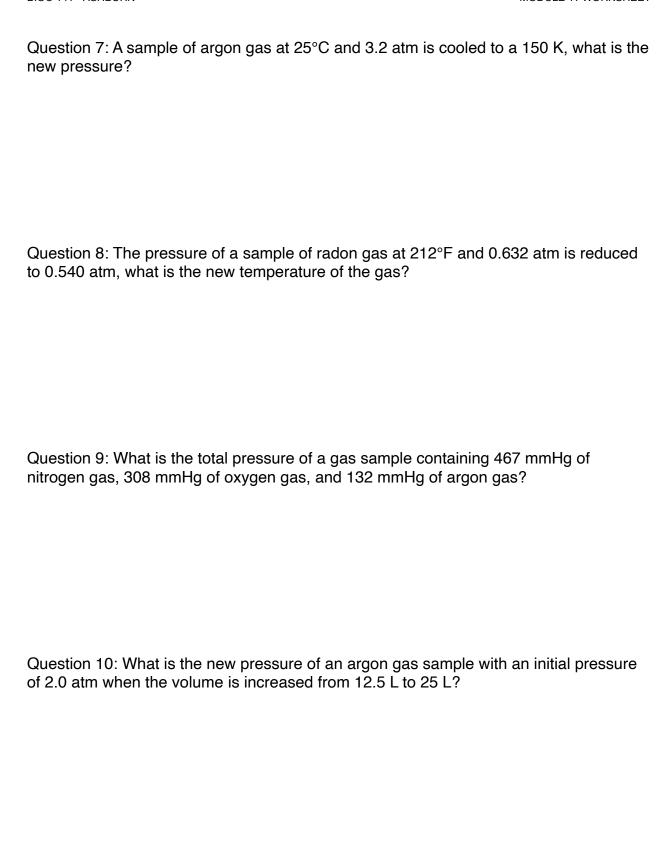
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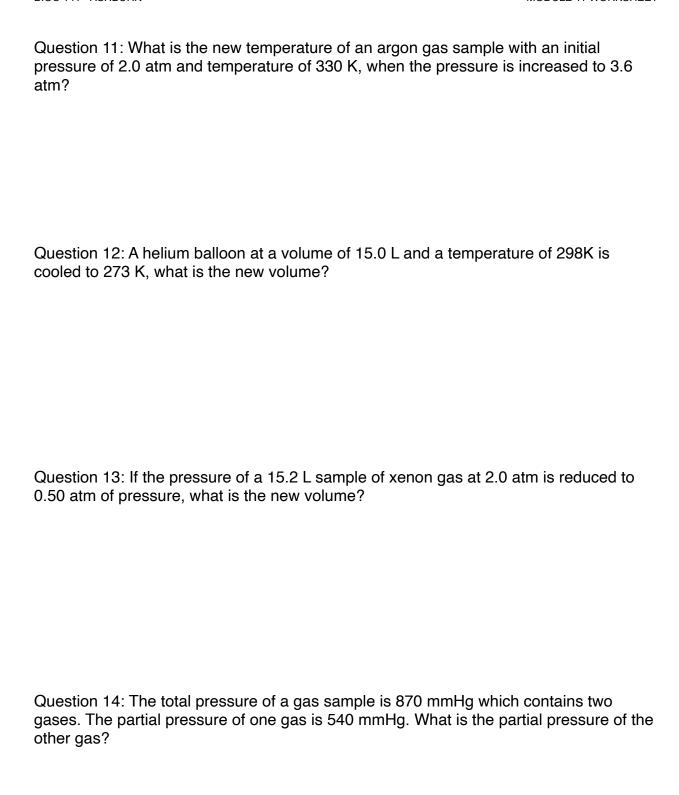
Question 1: Convert the following measurements of pressure into the desired units of pressure. I recommend using the **Exam 3 Handout** found in Module 15 for a listing of pressure relationships.

 A patient's systolic blood pressure is 128 mmHg, what is the pressure in units of atm?
 A car tire has a pressure of 35.0 psi, what is the pressure in units of mmHg?
 The atmospheric pressure at the top of Mt. Haleakala on Maui is 0.691 atm, what is the pressure in units of torr?
 Glaucoma can be diagnosed by measuring the intraocular pressure (IOP) inside the eye. A patient has an IOP of 23.2 mmHg, what is the value in kPa?
Question 2: Using your knowledge of Boyle's Law, Charles' Law, and Gay-Lussac's Law, predict what would happen to a sample of gas in each scenario.
• When pressure decreases, the volume would (increase or decrease).
• When temperature decreases, the volume would (increase or decrease).
When volume increases, the pressure would (increase or decrease).
When pressure increases, the temperature would (increase or decrease).
• When temperature increases, the volume would (increase or decrease).
When volume decreases, the temperature would (increase or decrease).
When pressure increases, the volume would (increase or decrease).

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Question 3: A sample of xenon gas has a volume of 3.0 L at a pressure of 8.0 atm. What is the volume of the gas at a pressure of 12.0 atm?	ıt
Question 4: A sample of carbon dioxide gas has a volume of 150 mL at a pressure of 0.25 atm. What is the pressure if the volume is changed to 30. mL?	
Question 5: A 25.0 L sample of neon gas at 96 K is heated to 220 K, what is the new volume?	
Question 6: A 2.60 L sample of chlorine gas at 100. K is compressed to a new volume of 1.42 L, what is the new temperature?	of





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Question 15: F	or each molecule below, list all the intermolecular forces it would exhibit.
C ₈ H ₁₈	
CH ₄ O	
Cl ₂	
HCI	
CH₅N	